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3 (Sem-5/CBCS) CHE HC 1

2024

**CHEMISTRY**

(Honours Core)

Paper : CHE-HC-5016

**(Organic Chemistry - IV)**

Full Marks : 60

Time : Three hours

***The figures in the margin indicate  
full marks for the questions.***

1. Answer the following questions :  $1 \times 7 = 7$
- (a) Adenosine is a \_\_\_\_.
  - (b) Adenosine 5'- monophosphate is a \_\_\_\_.
  - (c) Why  $\alpha$  - amino acids (except glycine) are optically active ?
  - (d) Give an example of Dipolar ion.

Contd.



- (e) Give an example of metalloenzyme.
- (f) Give an example of triacylglycerol.
- (g) Lauric acid is a \_\_\_\_ acid.

2. Answer the following questions :  $2 \times 4 = 8$

- (a) What are the *four* different bases present in DNA ?
- (b) Draw the structures of adenosine and 2'-deoxyadenosine.
- (c) Write down the equations for the reaction of glycine with  $\text{NaOH}_{(aq)}$  and  $\text{HCl}_{(aq)}$ .
- (d) What do you mean by functional group interchange (FGI) and functional group addition ?

3. Answer **any three** of the following :  $5 \times 3 = 15$

- (a) Explain the statement -  
"ATP is the carrier of Chemical Energy".

- (b) Indicate whether each functional group of the five heterocyclic bases in nucleic acids can function as a hydrogen bond acceptor, (A), a hydrogen bond donor (D), or both (D/A).

- (c) What do you mean by pI value of an amino acid ? Which amino acid has the lowest pI value and which amino acid has the highest pI value ? Give reasons.

- (d) What are the enzymes and co-enzymes ? Give examples. (*one for each*)

- (e) Define Saponification number and Iodine number. In what way these have proved useful in the analysis of oils and fats.

- (f) (i) Why are the carboxylic acid groups of the amino acids much more acidic ( $pK_a \sim 2$ ) than a carboxylic acid ( $pK_a \sim 4.76$ ) such as acetic acid.





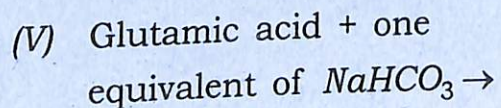
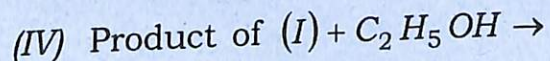
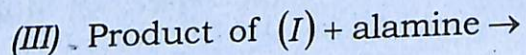
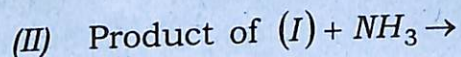
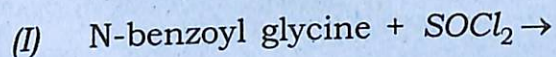
- (ii) Draw the form in which each of the following amino acids predominantly exists at physiological pH. (pH = 7.3)  
aspartic acid, glutamine, arginine, lysine, histidine, tyrosine

4. Answer **any three** of the following :

10×3=30

- (i) (a) What do you mean by a peptide bond ? Draw a structure of dipeptide by depicting the N-terminal and C-terminal amino acids. 2+3=5

- (b) Predict the products of the following reactions : 1×5=5



- (ii) (a) Write *one* method of each of synthesis of adenine and thymine.

- (b) Describe a method how the C-terminal residue of a polypeptide chain can be analysed.

- (c) Name *one* amino acid which is not found in  $\alpha$ -helix.

5+4+1=10

- (iii) Write short notes on the following :

3+3+4=10

- (a) Oxidation of food stuffs and cellular energy

- (b) Catabolism and anabolism

- (c) Metabolic pathways of carbohydrates

- (iv) (a) Write a method of synthesis of paracetamol

- (b) Mention *four* qualities that an antibiotic must possess.



(c) Mention *one* medicinal value of turmeric and neem.

(d) Name *two* useful drugs which are employed as antimalarials.

(e) Give a synthetic method for chloramphenicol.

2×5=10

(v) (a) Draw the structures of DNA and RNA.

(b) If one of the strands of DNA has the following sequence of bases running in the 5'→3' direction

5'-G-G-A-C-A-A-T-C-T-G-C-3'

What is the sequence of bases in the complementary strand? What are the forces that keep the *two* strands of DNA together?

5+5=10

(vi) Write short notes on : 2½×4=10

(a) Lipids

(b) Enzymes

(c) Nucleic Acids

(d) Polypeptides





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3 (Sem-5/CBCS) CHE HC 2

2024

**CHEMISTRY**

(Honours Core)

Paper : CHE-HC-5026

**(Physical Chemistry - V)**

Full Marks : 60

Time : Three hours

***The figures in the margin indicate  
full marks for the questions.***

1. Answer the following as directed :  $1 \times 7 = 7$ 
  - (a) State whether the following statement is True **or** False :  
State function  $\psi$  of a system must be an eigenfunction of the operator  $\hat{H}$ .
  - (b) Indicate whether the function  $\psi = e^{-x}$  is acceptable as wave function or not.  
 $(-\infty \leq x \leq \infty)$

Contd.



- (c) Show whether the operator  $\hat{O}$  in the equation  $\hat{O}\psi = \psi^2$  is linear or not.
- (d) State the rule of mutual exclusion in connection with Raman spectroscopy.
- (e) State the basic difference between line spectrum and continuous spectrum.
- (f) Calculate the energy per mole of light having wavelength of 8000Å.
- (g) What are photosensitizers ? Give *one* example.

2. Answer the following questions :  $2 \times 4 = 8$

- (a) Normalize the function :

$$\psi = \sin \frac{\pi x}{a}; 0 \leq x \leq a$$

- (b) It is required that the eigenfunction of an operator representing a physical quantity should be single-valued, continuous and quadratically integrable. State why the function should be single-valued and continuous.

- (c) What is the lowest vibrational energy in terms of oscillation frequency for a diatomic molecule undergoing simple harmonic motion ? Give the expression. What does it imply ?

- (d) The Stokes' lines are more intense than the anti-Stokes' lines. Explain.

3. Answer **any three** questions :  $5 \times 3 = 15$

- (a) (i) A certain system is described by

$$\text{the operator } \hat{A} = -\frac{d^2}{dx^2} + x^2.$$

Show that  $\psi = \left( x e^{-x^2/2} \right)$  is eigenfunction of  $\hat{A}$ .

What is the eigenvalue ?  $2+1=3$

- (ii) Calculate the spacing between the first two energy levels for an electron in a one-dimensional box of 1.0Å length. 2

- (b) Write how the molecular orbitals of a homonuclear diatomic molecule can be classified as  $\sigma$  and  $\pi$ . Which of these two is doubly degenerate and why ? What is the basis of classifying the MOs as  $g$  and  $u$  ?  $2+2+1=5$



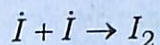
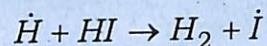
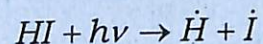
- (c) The first line in the pure rotational spectrum of  $HCl$  appears at  $21.18\text{ cm}^{-1}$ . Calculate the bond length of the molecule. Given atomic masses of  $H$  and  $Cl$  are  $1.008\text{ amu}$  and  $35.45\text{ amu}$  respectively.

$$(1\text{ amu} = 1.66 \times 10^{-27}\text{ kg}) \quad 5$$

- (d) (i) What do you mean by fingerprint region in IR spectroscopy? Explain with example.  $1+1=2$

- (ii) Write the quantum mechanical theory of Raman effect. What do you mean by Raman shift?  $2+1=3$

- (e) The decomposition of  $HI$  takes place by the following mechanism :



Find an expression for the rate of the reaction. Also find the quantum efficiency of the reaction.  $3+2=5$

4. Answer **any three** questions :  $10 \times 3 = 30$

- (a) (i) What is a Hermitian operator? Show that the eigenvalue of a Hermitian operator is real.

$$2+3=5$$

- (ii) Evaluate the expectation values of  $\langle \hat{x} \rangle$  and  $\langle \hat{p}_x \rangle$  for the ground state of the harmonic oscillator. Given normalized wave function

$$\psi_{(x)} = \left( \sqrt{\frac{a}{\pi}} \right) e^{-ax^2/2};$$

standard integral

$$\int_{-\infty}^{\infty} x^2 e^{-ax^2} dx = \left( \frac{1}{2a} \right) \left( \frac{\pi}{a} \right)^{1/2} \quad 5$$

- (b) (i) Write the general expressions for the magnitude and z-component of angular momentum. Write what you mean by space quantisation of angular momentum. Discuss with diagram the orientations of angular momentum of magnitude  $\sqrt{2}\hbar$  in presence of applied magnetic field in z-direction.

$$1+2+2=5$$



- (ii) Write what do you mean by radial distribution function. Find an expression for the radial distribution function. Draw the plot of radial distribution function against the radial distance from the nucleus for 1s orbital. State how this plot differs from the plot of square of the radial function against the radial distance.

1+2+1+1=5

- (c) (i) Discuss the LCAO-MO concept with respect to the hydrogen molecule ion. 5

- (ii) Why does  $He_2^+$  exist whereas  $He_2$  does not ? 2

- (iii) For a particle in a, one-dimensional box of length  $a$ , find the probability of finding the particle in the region  $0 \leq x \leq \frac{a}{4}$  in the ground state. 3

- (d) (i) Show that for a rotational spectrum of a diatomic molecule, the rotational quantum number (to the nearest integer value) for the maximum populated level is given

$$\text{by } J_{\max} = \sqrt{\frac{kT}{2hcB}} - \frac{1}{2}$$

5

- (ii) Considering diatomic molecule to be an anharmonic oscillator, deduce the energy expressions for the allowed vibrational transitions. Explain fundamental absorption and overtones. 5

- (e) (i) Using the Franck-Condon principle, explain why the intensities of the vibrational lines associated with electronic transitions differ. 3

- (ii) Define auxochrome. What do you mean by red shift and blue shift of absorption maximum ? 1+1=2



(iii) In the pure rotational spectra of  $^{14}\text{NO}$  and  $^{15}\text{NO}$ , the first lines appear at  $3.4\text{cm}^{-1}$  and

$3.2815\text{cm}^{-1}$  respectively. If the atomic mass of  $^{14}\text{N}$  and  $\text{O}$  are  $14.004\text{amu}$  and  $15.9994\text{amu}$ , respectively, find the atomic mass of  $^{15}\text{N}$ . 5

(f) (i) Write the mechanism of the  $\text{H}_2 - \text{Cl}_2$  photochemical reaction. Prove that the rate of formation of  $\text{HCl}$  is directly proportional to the intensity of the absorbed radiation.

2+3=5

(ii) In a certain photochemical reaction using  $464\text{nm}$  radiation, the incident light power was  $16.0\text{W}$  and the system absorbed 75% of the incident light. The quantum yield of the reaction was found to be 0.15. How many moles of the product were formed in the reaction in 100s? 5