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3 (Sem-1/CBCS) CHE HC 1

2022

**CHEMISTRY**

(Honours)

Paper : CHE-HC-1016

**(Inorganic Chemistry-I)**

Full Marks : 60

Time : Three hours

**The figures in the margin indicate full marks for the questions.**

1. Answer **any seven** questions from the following : 1×7=7
- (a) Which series of spectral lines of hydrogen atom falls in visible range of sunlight ?
- (b) What is an eigenfunction ?
- (c) What is the value of shielding constant for the valence electron of Li atom ?

Contd.

- (d) Arrange the halogens in decreasing order of their electronegativities.
- (e) Calculate the formal charge of P in  $\text{PH}_4^+$  ion.
- (f) How does bond multiplicity affect bond length?
- (g) Can the molecule  $\text{Be}_2$  exist?
- (h) An element can show +1, +3 and +5 oxidation states. In which state does the element can act both as an oxidising agent as well as a reducing agent?
- (i) What is reduction potential?
- (j) In which group of the periodic table will the element having atomic number 119 lie?
- (k) What is an operator?
- (l) Which group of elements of the periodic table are collectively known as chalcogen?

2. Answer **any four** questions from the following : 2×4=8

- (a) If the value of Rydberg constant for H atom is  $1.1 \times 10^7 \text{ m}^{-1}$ , what would be its value for  $\text{He}^+$  ion?
- (b) Define micro and macro particles in terms of 'action'.
- (c) Show that the probability of finding the electron in a hydrogen-like atom is independent of the angle  $\phi$ .
- (d) Discuss the trend in basic strength of the hydrides of group 15 elements.
- (e) Second electron affinity of the elements is positive. Why?
- (f) Ionic radii of  $\text{O}^{2-}$  ion is greater than  $\text{O}^-$  ion. Why?
- (g) With the help of a proper example explain that a non-polar molecule can have polar bonds.
- (h) With the help of a proper example explain disproportionation reaction.

3. Answer **any three** questions from the following :  $5 \times 3 = 15$

- (a) What do you mean by orbital quantum numbers? How are their values interrelated?  $3 + 2 = 5$
- (b) What are symmetric and antisymmetric wave functions? Explain Pauli's antisymmetry principle.  $2 + 3 = 5$
- (c) How can the ionisation energy values of the elements be applied in determining reducing power and reactivity order of the elements.  $2\frac{1}{2} \times 2 = 5$
- (d) How does electronegativity vary with bond order and hybridisation? With the help of electronegativity concept explain that  $HClO_3$  is a stronger acid than  $HClO$ .  $2 + 3 = 5$
- (e) Give Kapustinskii expression for lattice energy and identify the terms involved. What are its utilities?  $3 + 2 = 5$
- (f) With the help of VSEPR theory explain the shapes of  $XeF_2$  and  $IF_5$  molecules.  $2\frac{1}{2} \times 2 = 5$

(g) Discuss electrostatic theory of H-bonding. Give reason why ice floats over water.  $3 + 2 = 5$

(h) Derive Nernst equation for measuring EMF of cell.

4. Answer **any three** questions from the following :  $10 \times 3 = 30$

- (a) (i) With the help of Bohr's theory, explain the line spectra of H-atom.
- (ii) Explain Hund's rule of maximum multiplicity.
- (iii) Show that an s orbital has spherical shape.
- (b) What is effective nuclear charge? Explain the variation in effective nuclear charge of the elements along a period. Give Slater's rules for calculating screening constant.  $3 + 2 + 5 = 10$
- (c) (i) Explain Allred and Rochow's approach of electronegativity. What are its limitations?  $3 + 2 = 5$

(ii) What are radial nodes? Draw the radial probability distribution curve for  $4p$ ,  $4d$ ,  $3s$  and  $5f$  orbitals. 1+4=5

(d) Give Hittler and London approach of bonding in  $H_2$  molecule.

(e) What is polarisation in ionic compounds? Give Fajan's rules on polarisation. With the help of polarisation explain the solubility of silver halide in water.

3+4+3=10

(f) (i) Discuss electron sea model of metallic bonding and explain the metallic properties arising out of it. 4

(ii) Explain various types of hybridisation involving  $s$  and  $p$  orbitals with suitable examples.

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(g) (i) What is ionisation energy? Explain the factors on which it depends.

1+4=5

(ii) Discuss the crystal structure of zinc blende. 5

(h) (i) Draw the resonance structure of  $CO_3^{2-}$  and  $SO_4^{2-}$ . 2

(ii) What do you mean by well-behaved function? 2

(iii) Identify whether the following functions are eigenfunctions of  $d^2/dx^2$  :

(a)  $\cos kx$

(b)  $e^{ikx}$

Find the eigenvalue (if any).

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(iv) Draw the shapes of the orbitals for  $l=2$ . 3

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3 (Sem-1/CBCS) CHE HC 2

2022

**CHEMISTRY**

(Honours)

Paper : CHE-HC-1026

**(Physical Chemistry - I)**

Full Marks : 60

Time : Three hours

**The figures in the margin indicate full marks for the questions.**

**(Symbols used signify their usual meaning)**

1. Answer the following as directed :

**(any seven)**

1×7=7

- (a) Write the kinetic gas equation for an ideal gas.
- (b) Define Boyle temperature of a gas.
- (c) Higher the viscosity, slower is the rate of flow of a liquid at a given temperature. (State True or False)

Contd.

(d) State the total number of planes of symmetry in a cubic system.

(e) The planes which are absent in simple cubic crystal system is

(i) (100)

(ii) (200)

(iii) (110)

(iv) (111)

*(Choose the correct option)*

(f) Explain why  $Zn^{2+}$  ion is not precipitated as  $ZnS$ , when  $H_2S$  gas is passed through a  $ZnSO_4$  solution in presence of  $HCl$ .

(g) Identify the odd pairs, that will not act as buffer solution, from the following :

(i)  $NH_4Cl$  and  $NH_4OH$  solution

(ii)  $NaCl$  and  $NaOH$  solution

(iii)  $CH_3COONa$  and  $CH_3COOH$  solution

(iv)  $NaH_2PO_4$  and  $Na_2HPO_4$  solution.

*(Choose the correct option)*

(h) The compression factor for hydrogen gas is always greater than 1. Explain.

(i) How vapour pressure of a liquid is related to its boiling point ?

(j) What is a thermotropic liquid crystal ?

(k) State the law of rational indices in a crystal system.

(l) What is universal indicator in acid-base titration ?

2. Answer the following questions : **(any four)-**  
2×4=8

(a) The mean free path of gas molecules increases and number of collisions per unit time decreases with the decreasing temperature. Explain.

(b) Explain qualitatively the structure of liquid water.

(c) Determine the interplanar spacing between the (221) planes of a cubic lattice of edge length 450 pm.

(d) The degree of hydrolysis of  $NH_4Cl$  in 0.02 M aqueous solution at 298 K is  $5 \times 10^{-3}$ . If  $pK_b$  for  $NH_4OH$  at 298 K is 4.73, calculate  $pH$  of the solution.

(e) Value of Van der Waals radius for gaseous  $A_2$  molecule is  $2.0 \text{ \AA}$ . Calculate Van der Waals constant  $b$  for the gas.

(f) Explain why sodium hydroxide solution is not used to precipitate  $Al^{3+}$  as  $Al(OH)_3$  in Gr 3 of qualitative analysis of salt ?

(g) With the help of a suitable example explain what is impurity defect in crystal.

(h) What is rotating crystal method of observing diffraction in single crystal ?

3. Answer **any three** of the following questions :  $5 \times 3 = 15$

(a) Starting from the Van der Waals equation, find an expression for the Boyle temperature of a gas. Calculate Boyle temperature of  $CO_2$  gas. (Given for  $CO_2$  gas  $a = 0.3637 \text{ Nm}^4 \text{ mol}^{-2}$  and  $b = 42.8 \text{ m}^3 \text{ mol}^{-1}$ )  $3 + 2 = 5$

(b) Obtain relations to express critical constants in terms of Van der Waals constants 'a' and 'b'. Show that critical compressibility constant  $Z_c$  of one mole of a Van der Waals gas is 0.375.  $4 + 1 = 5$

(c) Define coefficient of viscosity. Give the SI unit of coefficient of viscosity. Give the theory of determination of coefficient of viscosity of a liquid by Ostwald viscometer method.  $1 + 1 + 3 = 5$

(d) What is meant by symmetry elements and symmetry operations? With the help of suitable examples, explain what are centre of symmetry and axis of symmetry.  $2 + 3 = 5$

(e) Give one example each of a strong acid and a weak acid. Explain the role of solvent in the ionization process of these acids.  $2 + 3 = 5$

(f) Using the expression for Maxwell distribution of speed, show that average kinetic energy of a gas molecule is given by  $\frac{3}{2}kT$ .  $5$

(g) Explain the following :

(i) Surface tension of water increases on addition of  $NaCl$ .  $2$

(ii) Cleansing action of detergent.  $3$

(h) Discuss the theory of pH metric titration between acetic acid solution and sodium hydroxide solution. Show the graphical variation between pH change with volume of sodium hydroxide added for the titration.  $4 + 1 = 5$

4. Answer **any three** of the following questions :  $10 \times 3 = 30$

(a) Define collision frequency. Obtain an expression for bimolecular collision frequency of a pure gas. Explain how collision diameter of a gas can be calculated from the measurement of coefficient of viscosity of the gas.  
 $1+5+4=10$

(b) (i) What is the virial equation of state of a gas ? Express the Van der Waals equation of state in the virial form.  $2+4=6$

(ii) Explain the principle of continuity of states. 4

(c) (i) Show the graphical variations for distribution functions for speeds with speeds of a gas at temperatures

$T_1, T_2$  and  $T_3$  ( $T_1 < T_2 < T_3$ ). 3

(ii) Define vapour pressure of a liquid at a given temperature. Explain a method of experimental determination of vapour pressure of a liquid. How vapour pressure is related to the boiling point of a liquid ?  $1+4+2=7$

(d) Explain the powder X-ray diffraction method of determination of a crystal structure. Explain how the lattice planes in sodium chloride are indexed.  
 $5+5=10$

(e) (i) What is meant by ionic product of water ? Show that  $pH = \frac{1}{2} pK_w$  for pure water. If  $K_w = 4.0 \times 10^{-14}$  for pure water at 317K, calculate  $pOH$ .  
 $1+2+2=5$

(ii) Define buffer capacity. Explain the role of buffer solution in qualitative analysis of salt.  $1+4=5$

(f) (i) A bulb of capacity  $1 \text{ dm}^3$  contains  $3.011 \times 10^{23}$  gaseous He atoms. Pressure exerted by these molecules is  $101.325 \text{ kPa}$ . Calculate the temperature and root mean square speed of the gas, assuming it to behave ideally under the given condition. 4

(ii) Derive an expression for coefficient of viscosity of a gas relating the mean free path of the gas. How the viscosity of a gas is influenced by the change of pressure ?  $5+1=6$



- (g) (i) Write the Berthelot equation of state for a real gas, explaining the terms involved in it. How this equation is different from the Van der Waals equation of state ?  
2+2=4
- (ii) Show that at low density, the Van der Waals equation and the Dieterici equation gives the same value of pressure of a gas. 4
- (iii) Compare average velocity, root mean square velocity and most probable velocities of a gas at a given temperature. 2
- (h) (i) Derive the Henderson equations for acid and basic buffer solutions. 5
- (ii) Calculate the change in  $pH$  when  $0.05 \text{ cm}^3$  of  $1 \text{ M NaOH}$  solution is added to one litre of buffer solution containing  $0.1 \text{ M}$  acetic acid and  $0.1 \text{ M}$  sodium acetate at  $300 \text{ K}$ . Given that  $K_a$  for acetic acid at  $300 \text{ K}$  is  $2.0 \times 10^{-5}$ . 5
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